

Ch 9 HW: See Asses 20,22,24,25,28-30,32 Rev Con  
Sec 9.3 & 9.4 57,58,60

obj: Apply rules for naming and writing formulas for binary molecular compounds and acids.

## Binary Molecular Compounds

- Compounds w/ 2 nonmetal elements

$H_2O$ ,  $CO_2$ ,  $SO_2$  are always

- Naming Binary Molecular Compounds.

\* The subscripts in the compound part of the name.

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- The subscript is represented by a prefix.

- Steps

1) Start w/ the prefix for the first element followed by the element's name.

\* Do not use mono if the subscript is one.

2) Write the prefix for the second element w/ the name of the element and an IDE suffix.

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$\text{SeO}_3$   
Selenium Trioxide

$\text{H}_2\text{O}$   
Dihydrogen Monoxide

$\text{N}_2\text{O}_5$   
Dinitrogen Pentoxide  
Penta-oxide

$\text{F}_2\text{Br}$   
Difluorine Mono-bromide

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- Writing Formulas

\* The prefix represents the subscript for that element.

Nitrogen Trioxide      Diphosphorus Heptaoxide

$\text{N}_2\text{O}_3$

$\text{P}_2\text{O}_7$

Carbon Tetra Chloride

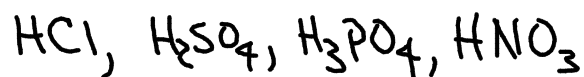
$\text{CCl}_4$

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## Acids

- Hydrogen compounds that will produce Hydrogen ions in water.

- Hydrogen is the Cation in the compound.



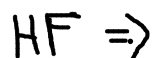
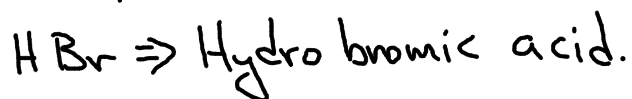
- The name of the acid depends on the name of the anion.

\* The suffix of the anion will determine the rule used to name the acid

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1) If the anion's suffix is IDE

then start w/ the word hydro  
name the anion w/ an ic suffix and  
end w/ the word acid.



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2) If the anion ends w/ an ITE  
Then name the anion w/ an ous suffix  
follow w/ the word acid.



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3) If the anion ends w/ ATE  
Then name the anion w/ an ic  
suffix and the word acid.



Carbonic Acid

Carbonate



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