

Ch 8

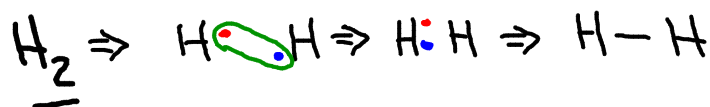
HW: Sec Asses 13-16,20,Rev Con 43,45,47

Sec 8.2

obj: Use Lewis dot structure show the formation of single, double and triple covalent bonds.

Covalent Bonds

- A force of attraction btw atoms that share valence electrons.

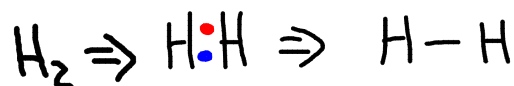


- 4 Types of Covalent Bonds

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1) Single Covalent Bonds

- 2 atoms share 1 pair of electrons btw them.
- 1 electron in the shared pair comes from each atom.



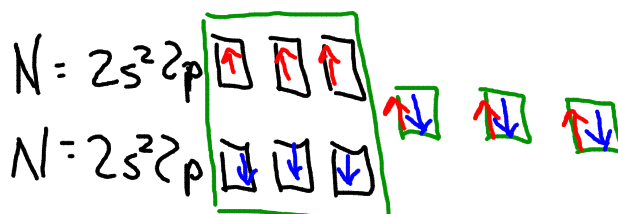
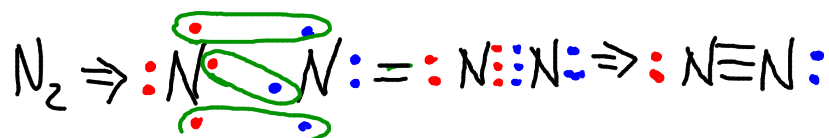
- Structural formula

- Represents the structure of the molecule and the bonds btw the atoms

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3) Triple Covalent Bond

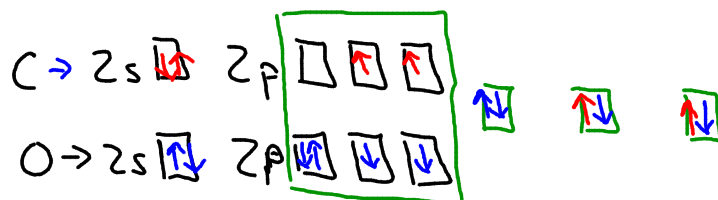
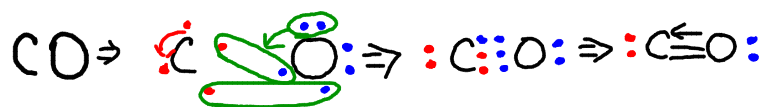
- 2 atoms share 3 pairs of electrons btw them.
- Each atom donates 3 electrons



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4) Coordinate Covalent Bond

- 1 atom donates both electrons btw the two atoms.



* Coordinate Covalent bonds act like any other covalent bond.

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