

Ch 13 HW: Sec Asses 8, 10-14 Rev Con 39, 40, 42, 43
Sec 13.2

obj: Describe the nature of a liquid in terms of the attractive forces between particles.

Properties of Liquids

- Liquid particles have weak attractive forces between them. (Intermolecular Forces)
 - * Reason why liquids take the shape of their container & have their own volume
- Vaporization - Change from a liquid to a gas
 - * Evaporation - Vaporization below the liquid's boiling Pt.
 - * Evaporation occurs at the surface of the liquid.
 - * As liquid particles evaporate they become vapor particles.

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- Vapor Pressure - Pressure exerted by the vapor particles from the liquid.
 - * Vapor pressure will produce an equilibrium in a contained liquid.
- Boiling - The rapid change of a liquid to a gas when the vapor pressure equals the external pressure acting on the liquid.
 - * Boiling Pt - Temperature where the liquid boils.
 - * Depends on the pressure acting on the liquid.
 - * Higher pressure higher boiling pt.

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* Normal Boiling Pt is the temperature @ which a Liquid boils when the external pressure is 1 atm.

- Both Liquids + gases are called fluids.

* Fluids can flow or be poured.

- Both Liquids + Solids are called condensed states of matter

* Can not be compressed into a smaller volume.

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