

Ch 11 HW: Sec 11,12 Rev Con 43,57

Sec 11.1A

obj: Write balanced chemical equations.

Balanced Chemical Equations

- Is a skeleton equation which also obeys the Law of Conservation of mass.

* Mass of the reactants equals the mass of the products.

* The number of atoms/ions for each element is the same on both sides of the equation.

11/20/02 8:19 AM

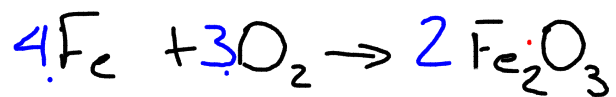
- Balancing Chemical Equations

1) Correct Skeleton Equation

2) Determine the # of atoms of each element on both the reactant and product side of the equation

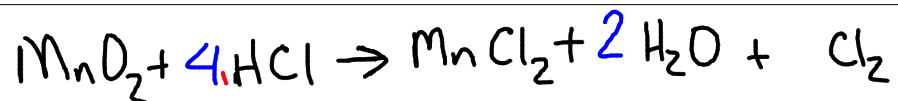
3) Use coefficients to balance the atoms for each element.

Nov 20 - 11:38 AM



	R	P
Fe	4	2
O	6	6

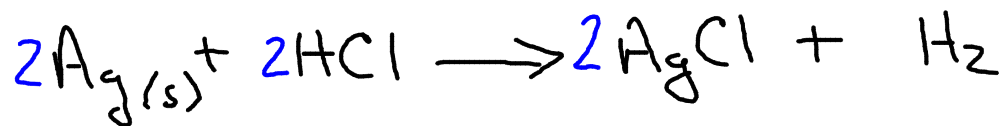
Nov 20 - 11:47 AM



	R	P
Mn	1	1
O	2	2
H	4	4
Cl	4	2

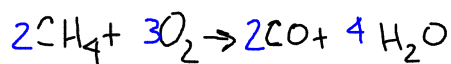
Nov 20 - 11:53 AM

Metallic Silver reacts w/ Hydrochloric acid producing Silver Chloride + Hydrogen gas.



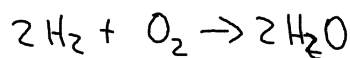
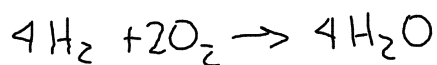
	R	P
Ag	2x	x2
H	2x	2
Cl	2x	x2

Nov 20 - 12:00 PM



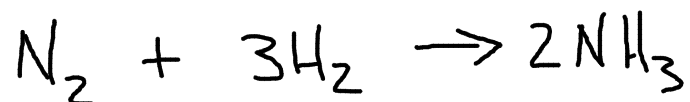
	R	P
C	2x	x2
H	8x	2x8
O	6x	2x6

* The coefficients in the balanced equations must always be written in smallest whole number ratios.



Nov 18 - 1:31 PM

Nitrogen + Hydrogen \rightarrow Ammonia



	R	P
N	2	1 2
H	6 2	3 6

Nov 9-11:02 AM

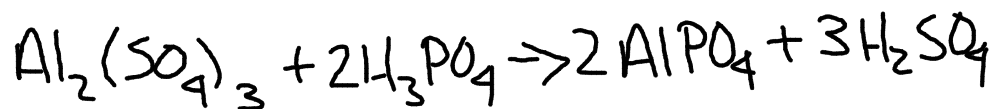
Butane (C_4H_{10}) + Oxygen \rightarrow Carbon Dioxide + Water



	R	P
C	8 4	1 48
H	20 10	2 1020
O	26 2	8 132426

Nov 9-11:04 AM

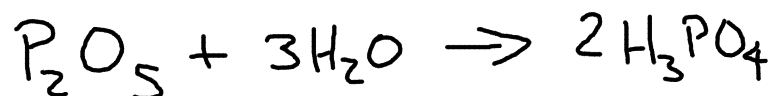
Aluminum Sulfate + Phosphoric Acid \rightarrow Aluminum Phosphate
+ Sulfuric Acid



	R	P
Al	2	1 2
S	3	1 3
O	20 6	8/ 16 20
H	6 3	26
P	24	1 2

Nov 9-11:05 AM

Diphosphorous Penta-oxide + Water \rightarrow Phosphoric Acid



	R	P
P	2	1 2
O	8 5	1 8
H	6 2	86

Nov 9-11:06 AM